

THREE-COMPONENT CARBON-CARBON BOND-FORMING REACTIONS FOR THE SYNTHESIS OF β -AMINO CARBONYL COMPOUNDS USING MALEIC ACID-CATALYZED ONE-POT MANNICH-TYPE REACTION IN ETHANOL

Malek Taher Maghsoodlou*¹, Seyyed Rasul Mousavi², Ali Roygar¹, Mojtaba Lashkari³

¹ Department of Chemistry, Faculty of Science, University of Sistan and Baluchestan, P.O. Box 98135-674, Zahedan, Iran

² Young Researchers and Elites club, Science and Research Branch, Islamic Azad University, Tehran, Iran

³ Faculty of Science, Velayat University, Iranshahr, Iran

*Corresponding author; E-mail: mt_maghsoodlou@chem.usb.ac.ir

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ABSTRACT. Herein, we described a one-pot, three-component and environmentally benign method for the synthesis of β -amino ketones based on the reaction of acetophenones, aromatic aldehydes, and aromatic amines using maleic acid as a green catalyst in ethanol as green solvent. The present methodology offers several advantages including good yields, mild reaction conditions and an inexpensive catalyst with a very easy work up. In addition, this method excludes the use of volatile organic solvents, tedious workup and column chromatographic purification of compounds, making the method greener, convenient and superior. All new compounds were well characterized by IR, ¹H and ¹³C NMR spectroscopy and physical data.

Keywords: three-component, maleic acid, Mannich reaction, one-pot, β -amino ketone, homogeneous catalyst.

INTRODUCTION

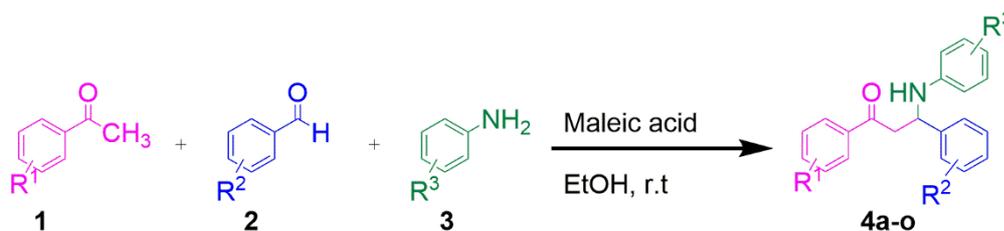
Carl Mannich first recognized the enormous significance of aminoalkylation of CH-acidic compounds, and then he extended the chemistry into a broad based synthetic procedure through systematic research (ARENDA *et al.*, 1998). Mannich reaction is one of the most important C-C bond forming reactions in organic synthesis for the preparation of secondary, tertiary amine derivatives and very useful compounds as building blocks in the synthesis of pharmaceuticals and natural products (KOBAYASHI and ISHITANI, 1999; IDHAYADHULLA *et al.*, 2011). The Mannich reaction is an important procedure for synthesizing of β -amino carbonyl compounds, which are significant synthetic intermediates for various pharmaceuticals and natural products (LOH *et al.*, 2000). The development of catalyst for synthetic methods has become an important research area, aiming to make the syntheses simpler, to save energy and to prevent toxicity in chemical processes. In recent years, the search for environmentally benign chemical processes or methodologies has received much attention (CHOUDHARY *et al.*,

2006; KURESHY *et al.*, 2010; MOUSAVI *et al.*, 2014; MOUSAVI *et al.*, 2015). Therefore, in chemistry “ideal synthesis” has been defined as one in which the target compound is generated in one step, in efficiency yield, from readily available and inexpensive starting materials in a resource-effective and environmentally acceptable process (SANDAROOS and DAMAVANDI, 2013). Hence, multi-component reactions (MCRs) have received much attention in the field of synthetic organic chemistry as well as medicinal chemistry, because the strategies of MCR offer significant advantages over conventional synthetic procedures (MADDILA *et al.*, 2016; AHMADI *et al.*, 2017). Therefore, projecting reactions in pursuit multi-bond formation in one operation is becoming one of the leading challenges in the field of green organic synthesis (MOUSAVI *et al.*, 2014). The approach of MCRs delivers numbers of advantages over conventional transformations, such as shorter reaction times, higher product yields, lower costs, an easy work-up process, and environmental being (MADDILA *et al.*, 2016D’SOUZA and MUELLER, 2007).

Organocatalysis has long existed as part of organic chemistry such as, acid or amine catalyzed Michael-additions, aldol condensation or Knoevenagel reactions (RANJBAR-KARIMI *et al.*, 2011; MOUSAVI and MAGHSOODLOU, 2014; MOUSAVI *et al.*, 2015; MOUSAVI and MAGHSOODLOU, 2015). Over the past decades, an extremely high output of results of different organocatalyzed reactions and transformations was heeded. Because, these transformations provide an extremely mild, selective access and operationally simple to required defined configured building blocks or natural products. Specifically, these include Mannich reactions, cycloadditions, aldol additions, conjugate additions etc (DÖMLING and UGI, 2000; MAHRWALD, 2013; SCHEFFLER and MAHRWALD, 2013;).

As noted above, β -amino carbonyl compounds are the main products of Mannich reaction and its derivatives are used for synthesizing peptides, amino alcohols, lactams and as precursors to optically active amino acids (HOWARD *et al.*, 2006). β -Amino carbonyl compounds forming via Mannich reaction are considerable and very important intermediates in the synthesis of natural products and pharmaceuticals (KOBAYASHI and ISHITANI, 1999; MÜLLER *et al.*, 1999; GUO *et al.*, 2007). Therefore, much attention has been drawn to the development of new synthetic methods to prepare these compounds. Traditionally, Mannich reactions have been achieved by using proton acids or Lewis acids as catalysts (LOH *et al.*, 2000; NOTZ *et al.*, 2001; KOBAYASHI and MANABE, 2002; WANG *et al.*, 2005; TERADA *et al.*, 2006; PHUKAN *et al.*, 2006; SAHOO *et al.*, 2006; YI and CAI, 2006; WANG *et al.*, 2007; LI *et al.*, 2007; WU *et al.*, 2007; KIDWAI *et al.*, 2009; YUE *et al.*, 2009; KURESHY *et al.*, 2010; SHARMA *et al.*, 2012; WANG *et al.*, 2012; ZHANG *et al.*, 2012; SAADATJOO *et al.*, 2017). However, despite recent spectacular advances, these procedures are not entirely satisfactory and suffer from some limitations including use of toxic reagents, atmosphere sensitive reagents or catalysts, a large amount of expensive and non-recoverable catalyst, requirement of special effort for catalyst preparation. Nevertheless, there is still a need for further efficient and, in particular, technically feasible methods in this field.

In continuation of our previous research toward the improvement of new green and convenient routes for the synthesis of various compounds using homogeneous catalysts (RANJBAR-KARIMI *et al.*, 2011; MOUSAVI *et al.*, 2013; MOUSAVI and MAGHSOODLOU, 2014; MOUSAVI *et al.*, 2014; MAGHSOODLOU *et al.*, 2015; MOUSAVI *et al.*, 2015), we report herein, the application of maleic acid as a homogeneous catalyst under mild conditions for efficient, convenient and facile green synthesis of various β -amino carbonyl derivatives through the one-pot, three-component reaction of aldehydes, amines and acetophenone in ethanol at room temperature. In addition, to the best of our knowledge, there are no reports on the use of maleic acid as a homogeneous catalyst for this conversion (Scheme 1).



Scheme 1. Synthesis of β -amino carbonyl derivatives.

As shown in Fig. 1, maleic acid is one of the isomers of butenedioic acid was used as an efficient, green and natural catalyst in this work for synthesizing β -amino ketones. In maleic acid the carboxylic acid groups are *cis*.

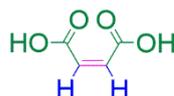


Figure 1. The structure of maleic acid.

MATERIALS AND METHODS

General

Melting point and IR spectra of all compounds were obtained on an Electrothermal 9100 apparatus and a JASCO FT/IR-460 plus spectrometer, respectively. ¹H and ¹³C NMR spectra of the new compounds were recorded on a Bruker DRX-400 Avance instrument in CDCl₃ as the solvent with TMS as internal standard at 400 and 100 MHz respectively. Elemental analyses for C, H, and N for the new compounds were performed using a Heraeus CHN-O-Rapid analyzer. All reagents were purchased from Merck (Darmstadt, Germany), Acros (Geel, Belgium) and Fluka (Buchs, Switzerland), and use without further purification.

General procedure for the synthesis of compounds 4

To a mixture of acetophenone **1** (1 mmol), aldehyde **2** (1 mmol) and amine **3** (1 mmol) in ethanol, 25 mol% maleic acid was added. The mixture was stirred at room temperature for an appropriate time (Table 4). The progress of the reaction was followed by thin-layer chromatography (TLC). After completion of the reaction, the precipitate was separated by simple filtration, and washed with ethanol (2 × 3 mL) and dried in the air to afford the title pure compounds. The identity of the known products was confirmed by comparison of their spectroscopic data and physical properties with those available in published articles. The novel products were identified by spectral data (i.e., IR, ¹H-NMR, ¹³C-NMR and CHN). The spectral data of the novel compounds as well as some representative compounds are described below.

3-((4-Chlorophenyl)amino)-3-(3-methoxyphenyl)-1-phenylpropan-1-one (**4n**)

White solid (59%), m.p. 155.5-156.5 °C; IR (KBr, $\nu_{\text{max}}/\text{cm}^{-1}$): 3383, 1669, 1603, 808. ¹H NMR (400 MHz, CDCl₃, δ / ppm): δ = 3.40 (dd, J = 16.4, 7.6 Hz, 1H, H-2), 3.49 (dd, J = 16.4, 5.2 Hz, 1H, H'-2), 3.97 (s, 3H, OCH₃), 4.59 (br s, 1H, NH), 4.92 (t, J = 5.6 Hz, 1H, H-3), 6.49 (d, J = 8.8 Hz, 2H, Ar-H), 6.87 (d, J = 8.8 Hz, 2H, Ar-H), 7.04 (d, J = 8.8 Hz, 2H, Ar-H), 7.34 (d, J = 8.4 Hz, 2H, Ar-H), 7.46 (t, J = 7.6 Hz, 2H, Ar-H), 7.58 (t, J = 7.6 Hz, 1H, Ar-H), 7.92 (d, J = 7.2 Hz, 2H, Ar-H). ¹³C NMR (100 MHz, CDCl₃, δ / ppm): δ = 46.2, 54.3,

55.2, 114.2, 114.9, 122.3, 127.3, 128.1, 128.7, 133.5, 134.3, 145.3, 197.5. Anal. Calcd for C₂₂H₂₀ClNO₂: C, 72.23; H, 5.51; N, 3.83. Found: C, 72.35; H, 5.58; N, 3.89.

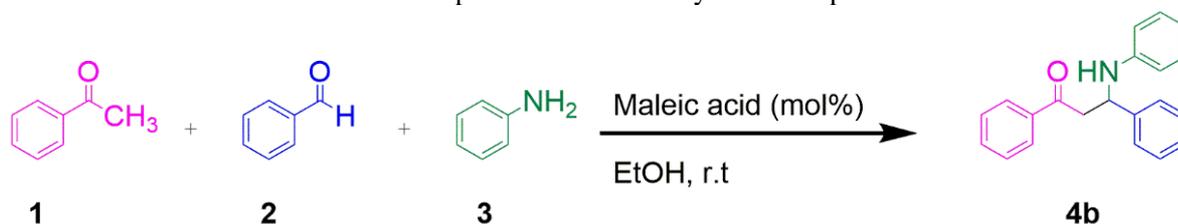
3-((4-Bromophenyl)amino)-3-(3-methoxyphenyl)-1-(p-tolyl)propan-1-one (4o)

White solid (45%), m.p. 151-152 °C; IR (KBr, $\nu_{\max}/\text{cm}^{-1}$): 3380, 1665, 1651, 503. ¹H NMR (400 MHz, CDCl₃, δ / ppm): δ = 2.42 (s, 3H, CH₃), 3.36 (dd, J = 16.0, 7.6 Hz, 1H, H-2), 3.45 (dd, J = 16.0, 5.2 Hz, 1H, H'-2), 3.97 (s, 3H, OCH₃), 4.63 (br s, 1H, NH), 4.89 (dd, J = 7.2, 5.2 Hz, 1H, H-3), 6.44 (d, J = 8.8 Hz, 2H, Ar-H), 6.87 (d, J = 8.8 Hz, 2H, Ar-H), 7.16 (d, J = 8.8 Hz, 2H, Ar-H), 7.26 (d, J = 8.0 Hz, 2H, Ar-H), 7.33 (d, J = 8.8 Hz, 2H, Ar-H), 7.82 (d, J = 8.4 Hz, 2H, Ar-H). ¹³C NMR (100 MHz, CDCl₃, δ / ppm): δ = 21.6, 46.1, 54.3, 55.2, 109.4, 114.2, 115.4, 127.3, 128.3, 129.4, 131.7, 134.1, 134.4, 144.4, 146.0, 158.8, 197.9. Anal. Calcd for C₂₃H₂₂BrNO₂: C, 65.10; H, 5.23; N, 3.30. Found: C, 65.23; H, 5.34; N, 3.37.

RESULTS AND DISCUSSION

Initially, we conducted the three-component reaction of acetophenone, benzaldehyde and aniline in the presence of varying amounts of maleic acid as catalyst in ethanol at ambient temperature conditions. The corresponding β -amino ketone was synthesized. The results were summarized in Table 1 and showed that the reaction using 25 mol% maleic acid at room temperature proceeded in highest yield (entry 4, Table 1). As shown in Table 1 when the reaction was performed in the absence of maleic acid, the reaction did not progress even after 48 h (entry 8).

Table 1. Reaction optimization for the synthesis of product **4b**.^a

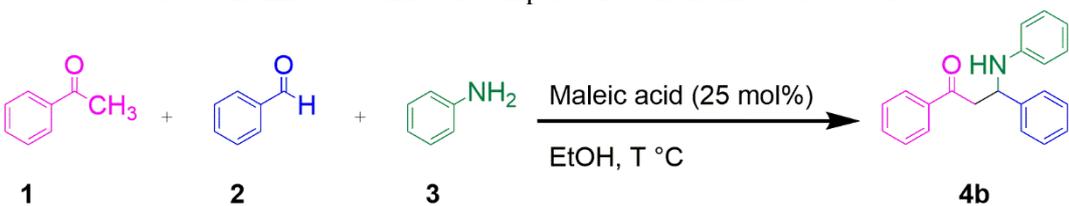


Entry	Catalyst (mol%)	Time (h)	Yield (%) ^b
1	10	18	55
2	15	16	63
3	20	17	79
4	25	14:30	83
5	30	14:15	77
6	35	14	77
7	40	13:30	78
8	-	48	-

^a The reaction was carried out using acetophenone (1 mmol), benzaldehyde (1 mmol) and aniline (2 mmol) in EtOH (5 mL) at room temperature.

^b Yield refers to the pure isolated product.

Next, to optimize the reaction temperature, the reaction of acetophenone, benzaldehyde and aniline was studied in the presence of 25 mol% maleic acid in EtOH at different temperatures. The results were summarized in Table 2 and showed that the reaction using 25 mol% maleic acid at room temperature proceeded in highest yield (entry 1, Table 2). Also as can be seen in Table 2, at the higher temperature did not increase the reaction yield (entries 2-4, Table 2).

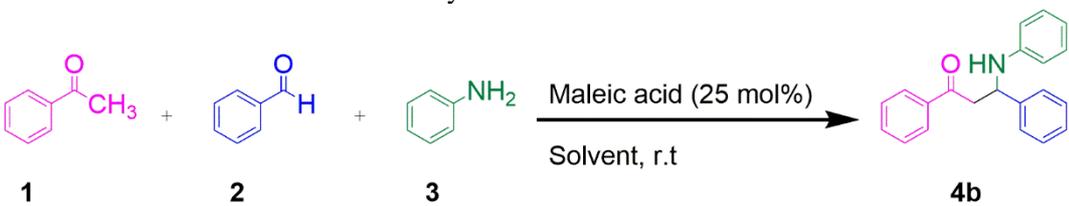
Table 2. Effect of different temperatures on the model reaction.^a


Entry	T (°C)	Time (h)	Yield (%) ^b
1	r.t	14:30	83
2	35	14:30	83
3	40	14	77
4	50	13	72

^a The reaction were carried out using acetophenone (1 mmol), benzaldehyde (1 mmol) and aniline (1 mmol) in the presence of 25 mol% maleic acid in EtOH (5 mL) at different temperatures.

^b Yield refers to the pure isolated products.

In the next step, the effect of different solvents on the model reaction was investigated at room temperature using 25 mol% of the catalyst. Various solvents like water, ethanol and the mixture of different ratio of ethanol/water, methanol and chloroform were screened to test the efficiency of our catalyst and the results are summarized in Table 3. The reaction in water and the ratio EtOH/H₂O (1:2) could not proceed to completion even after 72 h (entries 1 and 5, Table 3), chloroform as a solvent could not produce satisfactory results (entry 7, Table 3), using methanol, EtOH/H₂O (2:1) and EtOH/H₂O (1:1) as solvents gave average results with moderate product yield (entries 3, 4 ad 6, Table 3). As shown in Table 3, among the tested solvents, ethanol was found to be most efficient with excellent product yield in 14:30 h.

Table 3. Initial solvent effect studies for synthesis of β -amino ketone derivatives with 25 mol% catalyst at ambient conditions.^a


Entry	Solvent	Time (h)	Yield (%) ^b
1	H ₂ O	72	-
2	EtOH	14:30	83
3	EtOH/H ₂ O (2:1)	20	46
4	EtOH/H ₂ O (1:1)	30	46
5	EtOH/H ₂ O (1:2)	72	-
6	MeOH	17	46
7	CHCl ₃	30	Trace

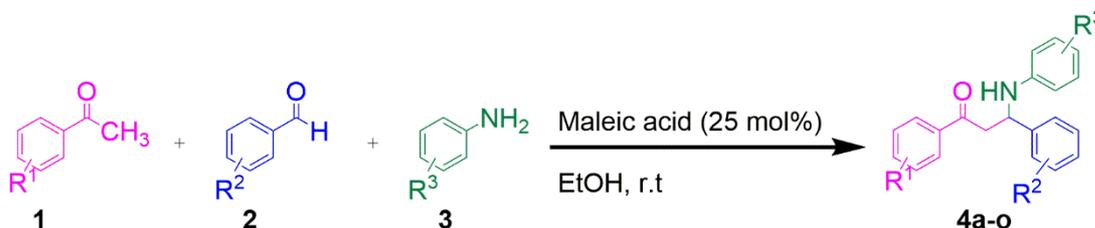
^a The reaction were carried out using acetophenone (1 mmol), benzaldehyde (1 mmol) and aniline (1 mmol) in the presence of 25 mol% maleic acid in various solvents at room temperature.

^b Isolated yield.

After the successful preparation of **4b**, we decided to introduce more diversity in the β -amino ketone scaffolds. With the optimized reaction conditions in hand, the scope of application of this three-component reaction was examined using different aldehydes, anilines and acetophenones as starting materials. Clearly, these reactions proceeded very cleanly under mild reaction conditions and no byproducts were observed. It is worthwhile to mention that

aromatic rings in starting materials with electron-donating and electron-withdrawing groups afforded the correspond products in moderate to good yields. The results are summarized in Table 4.

Table 4. Synthesis of β -amino carbonyls **4a-o**.



R ¹	R ²	R ³	Product	Time (h)	Yield (%) ^a	M.p (°C)	M.p Lit.
4-Me	H	H	4a	28	78	136-137	136-138
H	H	H	4b	14:30	83	167-169	167-169
H	4-Me	4-Me	4c	29	50	137-138	135-136
H	4-Me	H	4d	48	46	132-133	131-132
H	4-OMe	H	4e	12	50	146-147	147-149
H	H	4-Me	4f	12	57	167-168	168-170
H	H	4-Cl	4g	12:30	85	169-170	171-172
H	H	4-Br	4h	9	85	180-181	178-179
H	H	3-NO ₂	4i	27	65	142-143	140-142
H	4-Br	H	4j	15	70	131-133	128-130
H	4-Cl	4-Cl	4k	18	75	120-121	118-119
H	4-Cl	3-NO ₂	4l	16	50	130-131	132-133
4-Cl	3-NO ₂	H	4m	27	65	148-150	146-148
H	3-OMe	4-Cl	4n	18	59	1556-157	New ^b
4-Me	3-OMe	4-Br	4o	11	45	151-152	New ^b

^a Isolated yield.

^b The new compounds synthesized in this work.

The structure of all known synthesized compounds according to this procedure were characterized by comparison of the melting points and the analytical data with those reported for authentic samples. The structure of the new synthesized β -amino ketones **4** is in full agreement with ¹H NMR, ¹³C NMR and elemental analysis as illustrated below for a representative example (compound **4o**). The IR spectrum of **4o** displayed a sharp peak at 3380 cm⁻¹ was due to NH and one sharp strong absorption band was observed at approximately 1665 cm⁻¹ for C=O stretching. The ¹H NMR spectrum contained a characteristic single peak at δ = 2.42 ppm (s, 3H) for CH₃. A doublet of doublet was assigned at δ = 3.36 ppm (J = 16.0, 7.6 Hz) for one of the methylene protons (H-2). The another methylene proton of **4o** (H'-2) was observed as a doublet of doublet at δ = 3.45 ppm (J = 16.0, 5.2 Hz). A singlet was appeared at δ = 3.97 ppm (s, 3H) for OCH₃. Also, a characteristic single peak was observed at δ = 4.63 ppm (s, 1H) for NH. A doublet of doublet was assigned at δ = 4.89 ppm (J = 7.2, 5.2 Hz) for methine proton (H-3) (Fig. 2). The aromatic protons resonance were observed as several doublet at δ = 6.44-7.82 ppm. ¹³C NMR spectrum of **4o** was exhibited a peak at δ = 197.9 ppm for carbon of carbonyl group that is agreement with the structure of β -amino carbonyl **4o**.

In analogy with reported mechanisms (YUE *et al.*, 2009; LU and CAI, 2010; SHARMA *et al.*, 2012; MOUSAVI *et al.*, 2013), the following mechanism is proposed to explain the formation of product (Scheme 2). Maleic acid acts as an acid catalyst and facilitates the respective formation of imine (**6**) from aldehyde (**2**) and aniline (**3**) via removal of the water molecule. The next step of mechanism is the formation of enol (**7**) from acetophenone (**1**) in

the presence of catalyst. Next, inter-molecular Mannich type reaction occur between (6) and (7) to yield the product β -amino ketones (4).

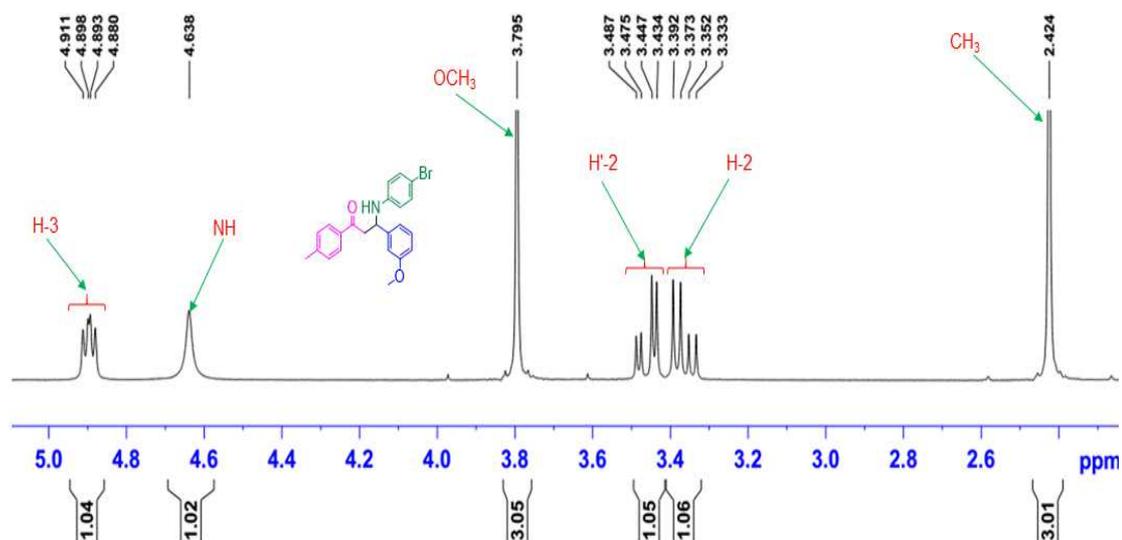
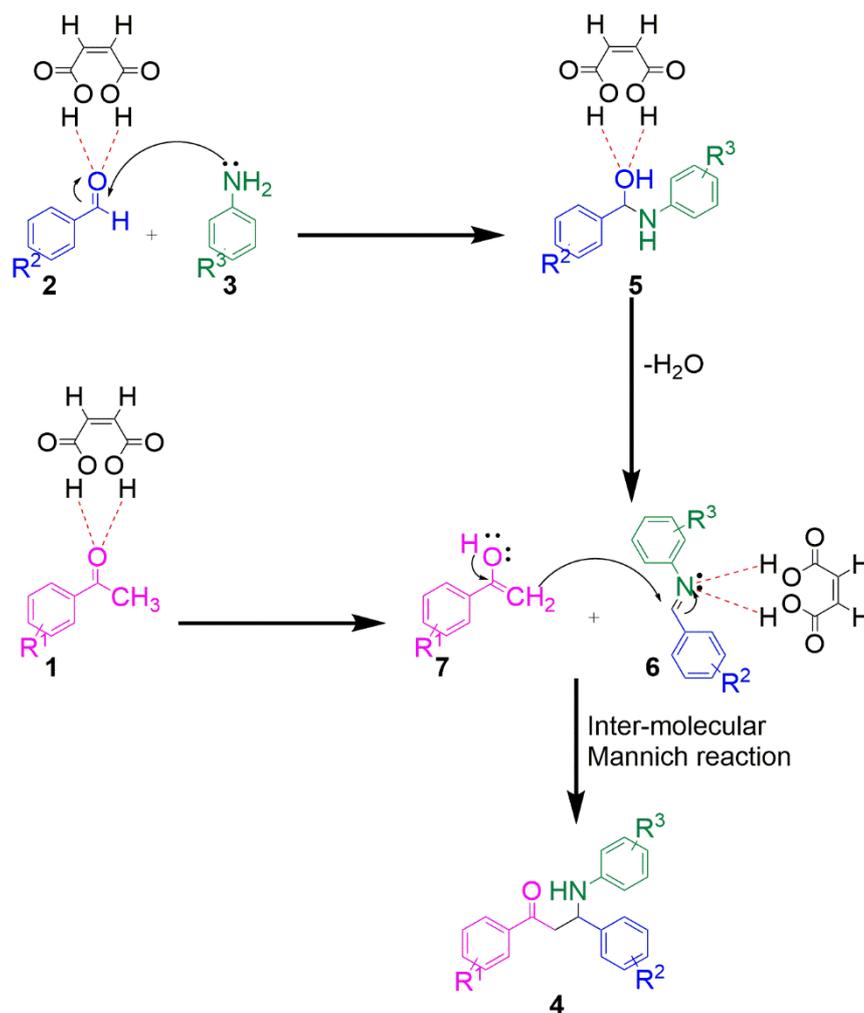


Figure 2. $^1\text{H-NMR}$ of **4o**.



Scheme 2. Plausible mechanism for one pot synthesis of β -amino ketone.

CONCLUSION

In summary, we have established effective and practical one-pot protocol for the synthesis of β -amino ketones by treatment of aromatic aldehydes, aromatic anilines with acetophenones in the presence of maleic acid as catalyst. The use of maleic acid in these MCRs has benefits such as clean reaction profiles, lack of side reactions, green properties, minimization of waste, a simple experimental procedure, high yields, simplicity of operation, and easy work-up. Moreover, all products were obtained through simple filtration with no need for column chromatography, which reduces the waste as well as environmental pollution.

Acknowledgments

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